

UNCLASSIFIED

AD. 4 6 4 2 6 0

DEFENSE DOCUMENTATION CENTER

FOR

SCIENTIFIC AND TECHNICAL INFORMATION

CAMERON STATION ALEXANDRIA, VIRGINIA



UNCLASSIFIED

NOTICE: When government or other drawings, specifications or other data are used for any purpose other than in connection with a definitely related government procurement operation, the U. S. Government thereby incurs no responsibility, nor any obligation whatsoever; and the fact that the Government may have formulated, furnished, or in any way supplied the said drawings, specifications, or other data is not to be regarded by implication or otherwise as in any manner licensing the holder or any other person or corporation, or conveying any rights or permission to manufacture, use or sell any patented invention that may in any way be related thereto.

PROOF BY LOWER BOUND METHODS THAT
NO SINGLY-EXCITED BOUND STATES EXIST IN H^-

by

William H. Millert
Department of Chemistry
Harvard University
Cambridge, Massachusetts

DDC

JUN 7 1965

1151A B

CATALOGED BY: DDC

464260

Variational calculations for the lowest $1S$ state in the hydrogen negative ion, H^- , showed¹ quite early that this state is bound. However, similar calculations^{2,3,4,5} for the lowest $3S$ state, the first excited $1S$ state, and the lowest ungerade $3P$ state have never produced an upper bound below the onset of the one-electron continuum. It has therefore been assumed that there are no singly-excited bound states in H^- . If calculated lower bounds for these states could be obtained and were found to lie above the onset of the continuum (-0.50 a.u. for these states), this would conclusively prove that these states cannot be bound.

The standard available lower bound methods, however, are not able to produce lower bounds above the beginning of the continuum. It is clear that the Temple⁶ and Weinstein⁷ methods do not apply here. Also, the methods of Bazley and Fox⁸, and the modification due to Gay⁹, cannot be applied because they require that the lower bound calculated from an l th order secular determinant be below E_{l+1}^0 , the $l+1$ st unperturbed eigenvalue. In H^- all these unperturbed eigenvalues E_{l+1}^0 lie below the onset of the one-electron continuum, so these methods can never produce a rigorous lower bound that lies above the beginning of this continuum.

AVAILABLE COPY WILL NOT PERMIT
FULLY LEGIBLE REPRODUCTION.
REPRODUCTION WILL BE MADE IF
REQUESTED BY USERS OF DDC.

464260

We have recently been able to extend a lower bound procedure¹⁰ that we presented earlier. As originally presented, our procedure is also limited to lower bounds lying below E_{l+1}^0 ; but we have now obtained a criterion that justifies the use of roots of the secular determinant above E_{l+1}^0 as rigorous lower bounds. In effect this criterion determines whether a given root of the determinant will increase or decrease if the order of the determinant is increased. We then reason as follows: if a state K is bound (i.e., its energy E_K is below -0.50 a.u.), then for some determinant sufficiently large, say $L \times L$, it will be true that E_K^L ¹¹ $< E_{L+1}^0$, and therefore E_K^L is a rigorous lower bound to E_K . However, if the K^{th} root of the determinant has only increased as the order of the determinant has been increased, then E_K^L for the smaller $l \times l$ determinant is also a lower bound to E_K -- even if $E_K^L > E_{l+1}^0$. We shall present the details justifying this extension of our lower bound method in a later publication.

The lower bounds we have obtained for H^- are presented in Table I. The number l given with the lower bound is the order of secular determinant used in each case. Note that even for the $1s^2$ state the "lower bounds" cannot at first glance be claimed as rigorous lower bounds since they are above E_{l+1}^0 . The criterion that we have developed, however, reveals that these roots must increase as the order of the determinant is increased, and they are therefore rigorous lower bounds.

We wish to emphasize that we are not claiming that there is a bound $1s2s$ 3S state in the one-electron continuum. The calculation says that the lowest bound 3S state must lie above -0.4871 -- and since there can

be no 3S bound states above -0.50 , no 3S bound states exist. In light of the work of O'Malley and Geltman¹² it seems clear that this numerical value is related to a resonance in the electron scattering off H atoms. However, the resonance will occur at such low energy (a few tenths of eV) that it is doubtful that it can be observed experimentally.

Our calculations give lower bounds for the $1s2s\ ^1S$ and $1s2p\ ^3P$ states that are above the lower bound for $1s2s\ ^3S$, and therefore these states are also not bound.

Another excited state of some interest is $2p^2\ ^3P$; it is the lowest 3P state and is bound in the He atom. The one-electron continuum for states of this symmetry begins at -0.1250 in H^- . E. Holþien has obtained³ an upper bound of -0.1243 , but we caution that though this is close to -0.1250 , it does not necessarily indicate that the state is "almost" bound. In fact, if a state is not bound, an unrestricted variational calculation will always converge to the onset of the continuum if a sufficiently flexible trial function is employed.^{2,4} However, the best (highest) lower bound that we have been able to obtain is -0.1260 (with an $k=5$ order determinant). We can only conclude from this that it is possible that $2p^2\ ^3P$ is bound in H^- . To prove that it is bound requires an upper bound that is below -0.1250 .

The author is happy to acknowledge his many helpful discussions with Professor E. Bright Wilson, Jr., concerning this problem.

TABLE I. Lower Bounds for H^- (atomic units, 1 a.u. = 27.2 ev)

<u>State</u>	<u>Upper Bound</u>	<u>Lower Bound (l)</u>	<u>E_{l+1}^o</u>	<u>Difference</u>
$1s^2 \ 1s$	-0.5278	-0.5623 (1)	-0.6250	0.0345
$1s^2 \ 1s$	-0.5278	-0.5536 (2)	-0.5556	0.0258
$1s2s \ 3s$	--	-0.4871 (1)	-0.5556	--

*This research was made possible by support extended Harvard University by the Office of Naval Research, Contract Nonr 1866 (14). Reproduction in whole or part is permitted for the U. S. Government.

NSF Predoctoral Fellow.

1. For an interesting history of the calculations of the $1s^2 \ ^1S$ state of H^- , see E. A. Hylleraas, *Astrophysica Norvegica* 9, 345 (1964).
2. G. K. Lowery and R. D. Present, *Astrophys. J.* 125, 611 (1957).
3. E. Holmøien, *Physica Norvegica* 1, 96 et. seq. (1961).
4. C. L. Pekeris, *Phys. Rev.* 112, 1649 (1958).
5. M. Machacek, F. C. Sanders, C. W. Scherr, *Phys. Rev.* 137, A1066 (1965).
6. G. Temple, *Proc. Roy. Soc. (London)* A119, 276 (1928).
7. D. H. Weinstein, *Proc. Nat. Acad. Sci.* 20, 529 (1934).
8. N. W. Bazley and D. W. Fox, *Phys. Rev.* 120, 144 (1961); *Phys. Rev.* 124, 483 (1961). Other references to this work are given in these papers.
9. J. G. Gay, *Phys. Rev.* 135, A1220 (1964).
10. W. H. Miller, *J. Chem. Phys.* 42, June 15, 1965.
11. E_K^L is the K^{th} root of the secular determinant of order L .
12. T. F. O'Malley and S. Geltman, *Phys. Rev.* 137, A1344 (1965).

UNCLASSIFIED
Security Classification

DOCUMENT CONTROL DATA - R&D

(Security classification of title, body of abstract and indexing annotation must be entered when the overall report is classified)

1. ORIGINATING ACTIVITY (Corporate author) Department of Chemistry Harvard University Cambridge, Massachusetts		2a. REPORT SECURITY CLASSIFICATION Unclassified	
2b. GROUP			
3. REPORT TITLE PROOF BY LOWER BOUND METHODS THAT NO SINGLY-EXCITED BOUND STATES EXIST IN H^-			
4. DESCRIPTIVE NOTES (Type of report and inclusive dates) Technical			
5. AUTHOR(S) (Last name, first name, initial) Miller, William H.			
6. REPORT DATE May 14, 1965		7a. TOTAL NO. OF PAGES 5	7b. NO. OF REFS 12
8a. CONTRACT OR GRANT NO. Nonr 1866		8a. ORIGINATOR'S REPORT NUMBER(S) --	
b. PROJECT NO. Task 14			
c.		8b. OTHER REPORT NO(S) (Any other numbers that may be assigned this report) --	
d.			
10. AVAILABILITY/LIMITATION NOTICES <i>See #10(1) next page</i> U. S. military agencies may obtain copies of this report directly from DDC. Other qualified users shall request through, Prof. E. B. Wilson, Harvard Univ.			
11. SUPPLEMENTARY NOTES -		12. SPONSORING MILITARY ACTIVITY Office of Naval Research Department of the Navy Washington, D. C.	
13. ABSTRACT <p>A previously presented lower bound method has been extended to make it possible to obtain lower bounds above the beginning of the one-electron continuum in a two-electron atom. The nature of this extension is briefly predicted and the method is applied to the eigenvalues of the hydrogen negative ion, H^-. It is found that the lowest bound 3S state must lie above -0.4871 a.u. -- and since there can be no bound states of this symmetry above -0.50 a.u. there are no bound 3S states in H^-. Also the lowest 3p_u ($1s2p$) state and the first excited $1s$ ($1s2s$) state have lower bounds above -0.50 a.u.</p> <p>The one-electron continuum for 3p_g states begins at -0.125 a.u. in H^-. The best lower bound obtained for the lowest state of this symmetry ($2p^2$) is -0.1260. Thus we cannot conclude that $2p^2$ 3p is not bound.</p>			

DD FORM 1473
1 JAN 64

UNCLASSIFIED
Security Classification

14. KEY WORDS	LINK A		LINK B		LINK C	
	ROLE	WT	ROLE	WT	ROLE	WT
Eigenvalues						
Lower Bounds						
Hydrogen Negative Ion						
Excited Electronic States, H ⁻						

INSTRUCTIONS

1. **ORIGINATING ACTIVITY:** Enter the name and address of the contractor, subcontractor, grantee, Department of Defense activity or other organization (*corporate author*) issuing the report.

2a. **REPORT SECURITY CLASSIFICATION:** Enter the overall security classification of the report. Indicate whether "Restricted Data" is included. Marking is to be in accordance with appropriate security regulations.

2b. **GROUP:** Automatic downgrading is specified in DoD Directive 5200.10 and Armed Forces Industrial Manual. Enter the group number. Also, when applicable, show that optional markings have been used for Group 3 and Group 4 as authorized.

3. **REPORT TITLE:** Enter the complete report title in all capital letters. Titles in all cases should be unclassified. If a meaningful title cannot be selected without classification, show title classification in all capitals in parenthesis immediately following the title.

4. **DESCRIPTIVE NOTES:** If appropriate, enter the type of report, e.g., interim, progress, summary, annual, or final. Give the inclusive dates when a specific reporting period is covered.

5. **AUTHOR(S):** Enter the name(s) of author(s) as shown on or in the report. Enter last name, first name, middle initial. If military, show rank and branch of service. The name of the principal author is an absolute minimum requirement.

6. **REPORT DATE:** Enter the date of the report as day, month, year, or month, year. If more than one date appears on the report, use date of publication.

7a. **TOTAL NUMBER OF PAGES:** The total page count should follow normal pagination procedures, i.e., enter the number of pages containing information.

7b. **NUMBER OF REFERENCES:** Enter the total number of references cited in the report.

8a. **CONTRACT OR GRANT NUMBER:** If appropriate, enter the applicable number of the contract or grant under which the report was written.

8b, 8c, & 8d. **PROJECT NUMBER:** Enter the appropriate military department identification, such as project number, subproject number, system numbers, task number, etc.

9a. **ORIGINATOR'S REPORT NUMBER(S):** Enter the official report number by which the document will be identified and controlled by the originating activity. This number must be unique to this report.

9b. **OTHER REPORT NUMBER(S):** If the report has been assigned any other report numbers (*either by the originator or by the sponsor*), also enter this number(s).

10. **AVAILABILITY/LIMITATION NOTICES:** Enter any limitations on further dissemination of the report, other than those

imposed by security classification, using standard statements such as:

- (1) "Qualified requesters may obtain copies of this report from DDC."
- (2) "Foreign announcement and dissemination of this report by DDC is not authorized."
- (3) "U. S. Government agencies may obtain copies of this report directly from DDC. Other qualified DDC users shall request through _____."
- (4) "U. S. military agencies may obtain copies of this report directly from DDC. Other qualified users shall request through _____."
- (5) "All distribution of this report is controlled. Qualified DDC users shall request through _____."

If the report has been furnished to the Office of Technical Services, Department of Commerce, for sale to the public, indicate this fact and enter the price, if known.

11. **SUPPLEMENTARY NOTES:** Use for additional explanatory notes.

12. **SPONSORING MILITARY ACTIVITY:** Enter the name of the departmental project office or laboratory sponsoring (paying for) the research and development. Include address.

13. **ABSTRACT:** Enter an abstract giving a brief and factual summary of the document indicative of the report, even though it may also appear elsewhere in the body of the technical report. If additional space is required, a continuation sheet shall be attached.

It is highly desirable that the abstract of classified reports be unclassified. Each paragraph of the abstract shall end with an indication of the military security classification of the information in the paragraph, represented as (TS), (S), (C), or (U).

There is no limitation on the length of the abstract. However, the suggested length is from 150 to 225 words.

14. **KEY WORDS:** Key words are technically meaningful terms or short phrases that characterize a report and may be used as index entries for cataloging the report. Key words must be selected so that no security classification is required. Identifiers, such as equipment model designation, trade name, military project code name, geographic location, may be used as key words but will be followed by an indication of technical content. The assignment of links, roles, and weights is optional.

Mr. Robert P. Clathin

DISTRIBUTION LIST FOR CONTRACT AF 33(615)-1635

No. of Copies

1. Air Force Materials Laboratory
Attn: MAAM (Lt. D. C. LaGrone)
Wright-Patterson AFB, Ohio 45433 3
2. Air Force Materials Laboratory
Attn: MAA (Mr. J. Teres)
Wright-Patterson AFB, Ohio 45433 1
3. Air Force Materials Laboratory
Attn: MAA (Technical Library)
Wright-Patterson AFB, Ohio 45433 1
4. Aerospace Industries Association
Attn: H. D. Moran
Technical Service Secretary
7660 Beverly Boulevard
Los Angeles, California 90036 1
5. Battelle Memorial Institute
Defense Metals Information Center
505 King Avenue
Columbus, Ohio 43201 1
6. Battelle Memorial Institute
Attn: Donald P. Moon
505 King Avenue
Columbus, Ohio 43201 1
7. Belfour Engineering Company
Technical Information Systems
Suttons Bay, Michigan 1
8. Boeing Transport Division
Attn: Ray E. Pearson
Staff Engineer - Titanium, P. O. Box 707
Renton, Washington 98055 1
9. Boeing Transport Division
Attn: David P. Donaldson
- Unit Chief of Structures Research
P. O. Box 707
Renton, Washington 98055 1
10. Chance Vought Corporation
Attn: G. A. Starr, Chief
Applied Research and Development
P. O. Box 5907
Dallas, Texas 75222 1

No. of Copies

- | | |
|--|---|
| 11. General Dynamics/Convair
Attn: C. W. Alesch, 6-143
P. O. Box 1950
San Diego, California 9211 | 1 |
| 12. Department of the Navy
Bureau of Weapons
Attn: R. Schmidt, RRMA-211
Washington 25, D. C. | 1 |
| 13. Douglas Aircraft Company, Inc.
Materials Research and Process Engineering
Attn: B. J. Alperin
3855 Lakewood Boulevard
Long Beach 1, California | 1 |
| 14. Douglas Missile & Space Division
Attn: Paul E. Denke
3000 Ocean Park Boulevard
Santa Monica, California | 1 |
| 15. General Dynamics/Ft. Worth
Attn: R. J. Patton
Fort Worth 1, Texas | 1 |
| 16. Harvey Engineering Laboratories
Test Engineering Branch
Attn: D. Q. Cole
19200 South Western Avenue
Torrance, California | 1 |
| 17. Lockheed California Company
Attn: M. A. Melcon
Structural Methods Department
P. O. Box 551
Burbank, California 91503 | 1 |
| 18. Lockheed-Georgia Company
Attn: W. T. Shuler
Department 72-01, Bldg. B-2, Zone 7
Marietta, Georgia 30061 | 1 |
| 19. Lockheed Missiles and Space Division
Attn: Dr. L. Schapiro
P. O. Box 504
Sunnyvale, California 94088 | 1 |
| 20. McDonnell Aircraft Corporation
Materials, Processes & Development Dept.
Attn: Howard J. Seigel
Lambert St. Louis Municipal Airport
St. Louis, Missouri | 1 |

No. of Copies

21. NASA Lewis Research Center
Attn: Wm. F. Brown, Jr.
Chief, Strength of Materials Branch
21000 Brookpark Boulevard
Cleveland, Ohio 44135
22. National Aeronautics and Space Administration
Scientific and Technical Information Division (SAK/DL) ✓
Head, Operations Section
P. O. Box 5700
Bethesda, Maryland 20014 1
23. North American Aviation
Attn: L. P. Spalding
Section Head, Materials
Los Angeles International Airport
Los Angeles, California 90009 1
24. North American Aviation
Attn: J. C. Joanides
Supervisor, Structures Research & Development
International Airport
Los Angeles, California 90009 1
25. Northrop, Norair Division
Materials and Processes Engineering
Attn: A. P. Binsacca
3101 W. Broadway
Hawthorne, California 1
26. Pratt & Whitney Aircraft Corp.
Attn: R. C. Huff
East Hartford, Connecticut 06108 1
27. Republic Aviation Corporation
Attn: F. W. Eversley, Jr.
Chief, Materials Applications
Farmingdale, L. I., New York 11735 1
28. Sikorsky Aircraft Company
Attn: W. G. Degnan
Stratford, Connecticut 1
29. Thompson Ramo Wooldridge, Inc.
Electromechanical Division
Attn: J. N. McCarthy
23555 Euclid Avenue
Cleveland, Ohio 44117 1

No. of Copies

- | | |
|--|---|
| 30. Titanium Metals Corporation
Attn: Ward W. Minkler
Manager of Technical Service
233 Broadway
New York, New York 10007 | 1 |
| 31. Titanium Metals Corporation of America
Attn: Robert R. Vogel
Metallurgical Engineer
6438 Corvette Street
Los Angeles, California 90022 | 1 |
| 32. U. S. Steel Corporation
Attn: J. J. Heger
522 Wm. Penn Place
Pittsburgh, Pennsylvania | 1 |
| 33. Wyman-Gordon Company
Attn: Paul J. Wisniewski,
Manager of Sales, Research &
Development & New Products
Worcester, Massachusetts 01601 | 1 |
| 34. U. S. Army Transportation Research Command
Attn: J. N. Daniel, Group Leader
Aeronautical Systems Group
Fort Eustis, Virginia 23604 | 1 |
| 35. Precision Forging Incorporated
Attn: F. D. Garz
600 St. Louis Street
Valley Park, Missouri | 1 |